

PERIODIC TABLE OF THE ELEMENTS

1A											8A										
1											18										
1	H																			2	
	1.008	2																		He	
																				4.003	
2	3	4																		10	
	Li	Be																		Ne	
	6.941	9.012																		20.18	
3	11	12	3B	4B	5B	6B	7B	8	8B	10	18	2B	13	14	15	16	17	18			
	Na	Mg											Al	Si	P	S	Cl	Ar			
	22.99	24.31											26.98	28.09	30.97	32.07	35.45	39.95			
4	19	20	21	22	23	24	25	26	27	28	29	30	31	32	33	34	35	36			
	K	Ca	Sc	Ti	V	Cr	Mn	Fe	Co	Ni	Cu	Zn	Ga	Ge	As	Se	Br	Kr			
	39.10	40.08	44.96	47.88	50.94	52.00	54.94	55.85	58.93	58.69	63.55	65.39	69.72	72.59	74.92	78.96	79.90	83.80			
5	37	38	39	40	41	42	43	44	45	46	47	48	49	50	51	52	53	54			
	Rb	Sr	Y	Zr	Nb	Mo	Tc	Ru	Rh	Pd	Ag	Cd	In	Sn	Sb	Te	I	Xe			
	85.47	87.62	88.91	91.22	92.91	95.94	(99)	101.1	102.9	106.4	107.9	112.4	114.8	118.7	121.8	127.6	126.9	131.3			
6	55	56	57	72	73	74	75	76	77	78	79	80	81	82	83	84	85	86			
	Cs	Ba	La	Hf	Ta	W	Re	Os	Ir	Pt	Au	Hg	Tl	Pb	Bi	Po	At	Rn			
	132.9	137.3	138.9	178.5	180.9	183.9	186.2	190.2	192.2	195.1	197.0	200.6	204.4	207.2	209.0	(209)	(210)	(222)			
7	87	88	89	104	105	106	107	108	109	110	111	112	113	114	115	116	117	118			
	Fr	Ra	Ac	Rf	Db	Sg	Bh	Hs	Mt	Ds	Rg	Cn	Nh	Fl	Mc	Lv	Ts	Og			
	(223)	226.0	227.0	(261)	(262)	(266)	(264)	(277)	(268)	(281)	(272)	(285)	(284)	(289)	(288)	(291)	(294)	(294)			

Lanthanides	58	59	60	61	62	63	64	65	66	67	68	69	70	71
	Ce	Pr	Nd	Pm	Sm	Eu	Gd	Tb	Dy	Ho	Er	Tm	Yb	Lu
	140.1	140.9	144.2	(145)	150.4	152.0	157.2	158.9	162.5	164.9	167.3	168.9	173.0	175.0
Actinides	90	91	92	93	94	95	96	97	98	99	100	101	102	103
	Th	Pa	U	Np	Pu	Am	Cm	Bk	Cf	Es	Fm	Md	No	Lr
	232.0	231.0	238.0	237.0	(244)	(243)	(247)	(247)	(251)	(252)	(257)	(258)	(259)	(260)

Fundamental Physical Constants

Avogadro's Number	$N_A = 6.02214 \times 10^{23} / \text{mol}$
Atomic Mass Unit	$\text{amu} = 1.66054 \times 10^{-27} \text{ kg}$
Charge of the Electron	$e = 1.60218 \times 10^{-19} \text{ C}$
Faraday Constant	$F = 9.64853 \times 10^4 \text{ C/mol}$
Mass of the Electron	$m_e = 9.10939 \times 10^{-31} \text{ kg}$
Mass of the Neutron	$m_n = 1.67493 \times 10^{-27} \text{ kg}$
Mass of the Proton	$m_p = 1.67262 \times 10^{-27} \text{ kg}$
Planck's Constant	$h = 6.62607 \times 10^{-34} \text{ J}\cdot\text{s}$
Speed of Light	$c = 2.99792 \times 10^8 \text{ m/s}$
Acceleration of Gravity	$g = 9.80665 \text{ m/s}^2$
Rydberg Constant	$R_H = 1.09677 \times 10^7 \text{ m}^{-1}$
Universal Gas Constant	$R = 8.31447 \text{ J/mol}\cdot\text{K}$ $R = 0.082058 \text{ L}\cdot\text{atm/mol}\cdot\text{K}$

Conversions and Relationships

Length	1 km = $1 \times 10^3 \text{ m} = 0.621 \text{ mile}$ 1 inch = 2.54 cm 1 ft = 12 in 1 pm = $1 \times 10^{-12} \text{ m} = 0.01 \text{ \AA}$
Mass	1 kg = $1 \times 10^3 \text{ g} = 2.205 \text{ lb}$ 1 metric ton = $1 \times 10^3 \text{ kg}$
Volume	1 dm ³ = $1 \times 10^{-3} \text{ m}^3 = 1 \text{ liter}$ 1 cm ³ = 1 mL 1 m ³ = 35.3 ft ³ 1 gallon = 3.785 liters
Energy	1 J = $1 \text{ kg}\cdot\text{m}^2/\text{s}^2 = 1 \text{ C}\cdot\text{V}$ 1 calorie = 4.184 J
Temperature	$T(\text{K}) = T(^{\circ}\text{C}) + 273.15$ $T(^{\circ}\text{C}) = (T(^{\circ}\text{F}) - 32)(5/9)$ H ₂ O: mp = 0°C and bp = 100°C
Pressure	1 Pa = $1 \text{ N/m}^2 = 1 \text{ kg/m}\cdot\text{s}^2$ 1 atm = $1.01325 \times 10^5 \text{ Pa}$ 1 atm = 760 torr = 760 mmHg
Math	$\pi = 3.1416$ $e = 2.7183$

Equations

$\Delta E = \Delta U = q + w$	$\Delta H = \Delta E + \Delta(PV)$	$q = mc\Delta T$	$w = -P_{\text{ext}}\Delta V$
$\Delta H_{\text{rxn}}^{\circ} = \sum \text{mol} \cdot \Delta H_f^{\circ}(\text{products}) - \sum \text{mol} \cdot \Delta H_f^{\circ}(\text{reactants})$			
$\Delta H_{\text{rxn}}^{\circ} = \sum \text{mol} \cdot BE(\text{bonds broken}) - \sum \text{mol} \cdot BE(\text{bonds formed})$			
$c = \lambda\nu$	$\Delta E = h\nu$	$\Delta E = \frac{hc}{\lambda}$	$\Delta E = -2.18 \times 10^{-18} \text{ J} \left(\frac{1}{n_f^2} - \frac{1}{n_i^2} \right)$
$M = \text{molar mass}$	$d = M/PRT$		$M = mRT/PV$
$PV = nRT$	$\frac{P_1V_1}{n_1T_1} = \frac{P_2V_2}{n_2T_2}$		$KE = \frac{3}{2}RT = \frac{1}{2}mv^2$
$F_c = \frac{kQ_1Q_2}{d^2}$	$v_{\text{rms}} = \sqrt{\frac{3RT}{M}}$		$\frac{\text{Rate}_A}{\text{Rate}_B} = \frac{\sqrt{M_B}}{\sqrt{M_A}}$
$P_A = X_A \cdot P_{\text{total}}$	$(P + n^2a/V^2)(V - nb) = nRT$		
$\ln\left(\frac{P_2}{P_1}\right) = \frac{-\Delta H_{\text{vap}}}{R} \left(\frac{1}{T_2} - \frac{1}{T_1} \right)$	$\ln\left(\frac{k_2}{k_1}\right) = \frac{-E_a}{R} \left(\frac{1}{T_2} - \frac{1}{T_1} \right)$		$k = Ae^{-E_a/RT}$
$P_{\text{solvent}} = X_{\text{solvent}}P^{\circ}_{\text{solvent}}$	$\Delta P = (X_{\text{solute}}P^{\circ}_{\text{solvent}})i$		$\Pi = (MRT)i$
$\Delta T_{\text{bp}} = (k_{\text{bp}}m)i$	$\Delta T_{\text{fp}} = (k_{\text{fp}}m)i$		$S_{\text{gas}} = k_{\text{H}}P_{\text{gas}}$
$[A]_t = -kt + [A]_0$	$\ln[A]_t = -kt + \ln[A]_0$		$\frac{1}{[A]_t} = kt + \frac{1}{[A]_0}$
$t_{1/2} = \frac{[A]_0}{2k}$	$t_{1/2} = \frac{\ln 2}{k}$		$t_{1/2} = \frac{1}{k[A]_0}$

Solubility Rules

- All common compounds of Group 1A ions and NH₄⁺ are soluble
- All common nitrates, acetates, and most perchlorates are soluble
- All common chlorides, bromides, and iodides are soluble, except those of Ag⁺, Pb²⁺, Cu⁺, and Hg₂²⁺. All common fluorides are soluble, except those of Pb²⁺ and Group 2A
- All common sulfates are soluble, except those of Ca²⁺, Sr²⁺, Ba²⁺, Ag⁺, and Pb²⁺
- All common metal hydroxides are insoluble, except those of Group 1A and the larger members of Group 2A (starting with Ca²⁺)
- All common carbonates and phosphates are insoluble, except those of Group 1A and NH₄⁺
- All common sulfides are insoluble, except those of Groups 1A, 2A, and NH₄⁺

Exam 2 – CHM 2045 – Fall 2020 – Study Review, Questions and Answers

Chapter 5-7 and ch 16

Question 1

10 pts

A sample of an ideal gas occupies 11.64 L at 464.6°C. Assuming that the pressure remains constant, what temperature (in °C) is needed to reduce the volume to 1.78 L?

Enter a number in degrees Celsius to 1 decimal places.

Question 2

10 pts

Not all pollution is due to human activity. Natural sources, including volcanoes, also contribute to air pollution. A scientist tries to generate a mixture of gases similar to those found in a volcano by introducing 15 g of water vapor, 5.4 g of SO₂, and 1.1 g of CO₂ into a 53 L vessel held at 136 °C.

Calculate the total pressure (in atm) of the mixture of gases. Enter a number to 4 decimal places.

Question 3

10 pts

An unknown liquid is vaporized in a 263 mL flask by immersion in a water bath at 104 °C. The barometric pressure is 753 torr. If the mass of the vapor retained in the flask is 1.618 g, what is its molar mass of the unknown liquid? The volume of the unknown liquid is negligible.

Enter your answer numerically (without units) to two decimal places.

Question 4

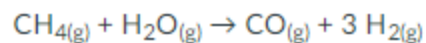
5 pts

Assuming ideal gas behavior, which of the following gases would have the lowest density at standard temperature and pressure?

 Kr CO₂ SF₆ CF₂Cl₂ OF₂**Question 5**

10 pts

Hydrogen gas (a potential future fuel) can be formed by the reaction of methane with water according to the equation:



In a particular reaction, 27.4 L of methane gas (measured at a pressure of 659 torr and a temperature of 24.2 °C) mixes with 25.9 L of water vapor (measured at a pressure of 686 torr and a temperature of 26.8 °C). The reaction produces 16.2 L of hydrogen gas at STP.

What is the percent yield of the reaction?
(round your answer up to one decimal place)

Question 6**5 pts**

At which pressure would H_2 be expected to behave more ideally?

- 100 atm
- 500 atm
- 1000 atm
- 1 atm
- 300 atm

Question 7**10 pts**

A system suffers an increase in internal energy of 115.2 J and at the same time the gas in the system is compressed from 352 mL to 294 mL by a constant external pressure of 4.8 atm. What is the heat change of the system?

$$101.3 \text{ J} = 1 \text{ L/atm.}$$

Report your answer with the correct sign (+/−) to 1 decimal place.

Question 8

10 pts

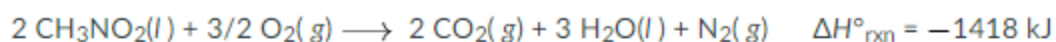
When 12.1 g KBr is dissolved in enough water to give 191 g of the solution in a coffee-cup calorimeter, the temperature drops from 53.1 °C to 22.9 °C. If the solution has a specific heat capacity which is essentially the same as water, 4.184 J/g·°C, calculate the enthalpy change in J per gram of KBr dissolved.

round your answer up to an integer number (to 0 decimal places)

Question 9

10 pts

Nitromethane (CH₃NO₂) burns in air to produce significant amounts of heat.



How much heat is produced by the complete reaction of 0.313 kg of nitromethane?

Enter your answer numerically up to 2 decimal places and in terms of kJ (do not use scientific notation). Enter a positive number.

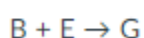
Question 10

10 pts

Given the series of reactions with known changes of enthalpy:



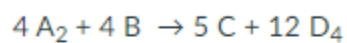
Figure out change of enthalpy (ΔH) of the reaction:



Write your answer in term of kJ/mol (number only) to 1 decimal place.

Question 11**10 pts**

Find the standard enthalpy of the schematic reaction (ΔH_{rxn}°) described by the equation:



Standard enthalpies of formation of reactants and products are (in kJ/mol)

$\Delta H_f^{\circ}(A_2)$	$\Delta H_f^{\circ}(B)$	$\Delta H_f^{\circ}(C)$	$\Delta H_f^{\circ}(D_4)$
164	91	357	-72

Write your answer in terms of kJ (number only) to 0 decimal places.

Question 12

10 pts

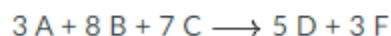
In a bomb calorimeter compartment surrounded by 1.047 kg of water, the combustion of 1.41 g of benzene, $C_6H_6(l)$, raised the temperature of the water from 23.640 °C to 32.692 °C. Find the heat capacity of the calorimeter in $kJ/^\circ C$. ΔE for the reaction is $-2,680.3328$ kJ/mol of $C_6H_6(l)$. The specific heat of water is 4.184 $J/g \cdot ^\circ C$.

- 965 $kJ/^\circ C$
- 0.874 $kJ/^\circ C$
- 1.056 $kJ/^\circ C$
- 652 $kJ/^\circ C$
- 0.965 $kJ/^\circ C$

Question 13

8 pts

Express the rate of this reaction in terms of the change in concentration of F (assume all are gases):



When C is decreasing at 1.984 M/s, how fast is F increasing?

Enter your answer without units to 3 decimal places.

Question 14

7 pts

Consider the reaction: $3 A + 5 B \longrightarrow 3 C$

If the rate of consumption of A is 0.924 M/s, what is the rate of reaction?

Enter a positive number without units to 3 decimal places.

Question 15**10 pts**

Given the individual reaction orders for all substances and the overall reaction order from the following rate law: $\text{rate} = k [A] [B]^2$

By what factor does the rate change if each of the following changes occur?

[Select] 0.5, 2, 0.25, 1, 4, 0.125, 8 [] is doubled; [2]

[Select] 0.5, 2, 0.25, 1, 4, 0.125, 8 [] is halved; [0.25]

[Select] 0.5, 2, 0.25, 1, 4, 0.125, 8 [] is doubled and [B] is halved? [0.5]

Question 16

10 pts

For the reaction $2 \text{NO} + \text{O}_2 \rightarrow 2 \text{NO}_2$, initial rate data are:

Experiment	[NO], M	[O ₂], M	Initial Rate, M/s
1	0.030	0.060	2.5×10^{-5}
2	0.030	0.020	2.5×10^{-5}
3	0.060	0.060	10.0×10^{-5}

The rate law is $\text{Rate} = k[\text{NO}]^x[\text{O}_2]^y$. What are the values of x and y respectively?

x = 2, y = 2

x = 0, y = 2

x = 2, y = 1

x = 2, y = 0

x = 1, y = 1

Question 17

10 pts

A decomposition reaction has a rate constant of $1.458 \times 10^{-3} \text{ 1/yr}$. How long does it take for reactant concentration to reach 39.8% of its original value.

Enter how many years (number only) to 0 decimal places.

Question 18

10 pts

For the reaction: $A(g) \rightarrow B(g) + C(g)$, the rate law is: $\text{rate} = k[A]$

At 300 K, the half-life is 345 seconds and the activation energy is 103.3 kJ/mol. What is the rate constant at 442 K?

Report your answer without units to 0 decimal places.

Question 19

5 pts

According to the Bohr model, an electron transition from $n = 5$ to $n = 3$ in a hydrogen atom represents emission of a longer wavelength than that of a transition from $n = 3$ to $n = 1$

 True False**Question 20**

8 pts

The frequency of a certain light is 5.2×10^{14} Hz, what is the wavelength in nm?

Enter a number only, in nm to 1 decimal place.

Question 21**10 pts**

A laser emits at a wavelength of 473 nm. All of its output energy is absorbed in a detector that measures a total energy of 1.32×10^{-16} J. How many photons are being emitted by the laser?

Enter the number of photons to 0 decimal places.

Question 22**10 pts**

How many orbitals in an atom can have the following quantum number designations. Enter a numerical value only (e.g. 0, 1, 2, 3 ...)

Designation	Number of Orbitals
$n = 3$	<input type="text" value="9"/>
$n = 4, l = 3$	<input type="text" value="7"/>

Question 23**2 pts**

Scratch paper: I am in the process of making very small pieces out of my scratch paper (aka confetti) and showing that process to the camera.

On my honor, I have neither given nor received unauthorized aid in doing this assignment.

 True False