

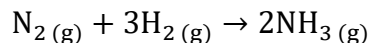
Spring 2024 CHM 2046 Exam 1 Review

The material covered in this review is from Chapters 16-19

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Chapter 16: Kinetics

1. Ammonia is generated on an industrial scale using the Haber-Bosch process. The reaction is shown below:



Find the rate law, individual, and overall reaction orders and the average value of k for the reaction.

Experiment	Initial rate (mol/L*s)	Initial [N ₂] (mol/L)	Initial [H ₂] (mol/L)
1	1.9×10^{-12}	0.0113	0.0011
2	1.7×10^{-11}	0.0220	0.0033
3	9.3×10^{-12}	0.0550	0.0011
4	4.9×10^{-11}	0.0220	0.0056

Rate Law:

N₂ Order:

H₂ Order

Overall Reaction Order:

Average Value of k

2. H₂O₂ decomposes into H₂ and O₂ in a first order reaction. If the initial concentration is 4.38 M, the final concentration is 2.91 M, and the decomposition takes place over 10 minutes, what is k? Using the calculated k, how long will it take to decompose 25% of the initial amount?

- a. k=0.035/min; 7 minutes
- b. k=0.041/min; 7 minutes
- c. k=0.035/min; 10 minutes
- d. k=0.041/min; 10 minutes
- e. k=0.059/min; 7 minutes
- f. k=0.059/min; 10 minutes

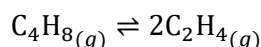
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3. Which of the following statements are true regarding exothermic reactions?

- | | | |
|-------------------------|---|--|
| I. Heat is absorbed | V. Heat and enthalpy will be on the same side of the equation | VII. The energy of the reactants is higher than the products |
| II. Heat is released | | |
| III. Heat is a reactant | VI. Heat and enthalpy will be on opposite sides of the equation | VIII. The energy of the reactants is lower than the products |
| IV. Heat is a product | | |

- I, III, V, VII
- II, IV, V, VII
- I, II, III, IV
- V, VI, VII, VIII
- I, IV, V, VIII
- II, III, V, VII

4. Cyclobutane decomposes in a first order reaction shown below.



Given that the initial concentration of C_4H_8 is 5M and the final concentration is 0.06M after 0.05 seconds, what is the rate constant and the expected rate law?

- 65 s^{-1} ; rate = $k[\text{C}_2\text{H}_4]^2$
- 88 s^{-1} ; rate = $k[\text{C}_4\text{H}_8]$
- 92 s^{-1} ; rate = $k[\text{C}_4\text{H}_8]$
- 88 s^{-1} ; rate = $k[\text{C}_2\text{H}_4]^2$
- 65 s^{-1} ; rate = $k[\text{C}_4\text{H}_8]$
- 92 s^{-1} ; rate = $k[\text{C}_2\text{H}_4]^2$

5. Which of the following statements are true regarding catalysts?

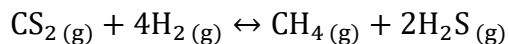
- | | | |
|---|--------------------------------------|--|
| I. Catalysts cause products to form slower | IV. Catalysts are not reformed | VII. Catalysts affect reaction rate; it increases |
| II. Catalysts cause products to form faster | V. Catalysts lower activation energy | VIII. Catalysts affect reaction rate; it decreases |
| III. Catalysts increase activation energy | VI. Catalysts are reformed | IX. Catalysts don't affect reaction rate |

- I, III, VI, IX
- II, III, VI, IX
- II, V, VI, VII
- I, IV, VI, VIII
- II, IV, VI, VII

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Chapter 17: Equilibrium

1. Given the following chemical reaction, calculate the K_c given that the K_p is 0.28 at 900°C.



- a. 7.5×10^{-5}
- b. 8.1×10^{-2}
- c. 3.6×10^{-3}
- d. 3.0×10^{-5}
- e. 2.9×10^{-4}

2. Which of the following statements regarding Q and K are true?

I. If $K > Q$, then the reaction proceeds to the right

II. If $K = Q$, then the reaction is at equilibrium

III. If the reaction proceeds to the right, it will create more products

IV. If the reaction proceeds to the left, it will create more products

V. If $K > Q$, then the reaction proceeds to the left

VI. If $K < Q$, then the reaction proceeds to the left

VII. If $K = Q$, then the reaction proceeds to the right

VIII. If $K < Q$, then the reaction is at equilibrium

- a. II, III, V, VIII
- b. I, II, III, VI
- c. IV, V, VI, VII
- d. VI, VII, VIII
- e. I, III, VI, VIII

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3. Fill in the table summarizing the effects of Le Chatelier's Principle.

Change	Effect on Equilibrium (Left or Right)	Effect on the value of K (Equilibrium Constant)
Increase [reactant]		
Increase [product]		
Decrease [reactant]		
Decrease [product]		
Increase pressure		
Increase volume		
Decrease pressure		
Decrease volume		
Increase pressure (inert gas)		
Increase temperature		
Decrease temperature		
Add catalyst		

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Chapter 18: Acid-Base Equilibria

1. Which of the following statements regarding acids, bases, and K_a s is true?

- | | |
|---|---|
| I. The stronger the acid, the larger the K_a , the larger the pK_a | VI. K_w , K_a , and K_b are related to each other in the equation $K_w = K_a * K_b$ |
| II. The stronger the acid, the larger the K_a , the smaller the pK_a | VII. The equilibrium of an acid base reaction goes from the stronger acid to the weaker acid |
| III. The weaker the acid, the lower the concentration of H_3O^+ , the larger the pK_a | VIII. The equilibrium of an acid base reaction goes from the weaker acid to the stronger acid |
| IV. The larger the pK_a , the smaller the K_a | IX. If the reaction proceeds to the right, $K_c > 1$. |
| V. A strong acid is a weak base | |
- a. I, VIII
 b. All but I, VIII
 c. II, III, VII, VIII
 d. IV, V, VII, VIII
 e. I, III, IV, VII

2. Which of the following statements regarding pH is true?

- | | |
|---|-------------------------------------|
| I. Acidic solutions have a higher concentration of OH^- | IV. $K_w = \frac{[H_3O^+]}{[OH^-]}$ |
| II. Basic solutions have a higher concentration of OH^- | V. $K_w = [H_3O^+] * [OH^-]$ |
| III. A neutral solution has an equal concentration of H_3O^+ and OH^- | VI. $pH + pOH = 14$ |
| | VII. $pH - pOH = 14$ |
- a. I, II, IV, VII
 b. II, III, V, VI
 c. III, IV, V, VII
 d. II, IV, VI

3. If an unknown weak acid is 0.798% dissociated in a 2.15M solution. What is the K_a of the acid, the pK_a , and the identity of the acid?

- a. $2.46 * 10^{-3}$, 10.5, Formic acid
 b. $5.12 * 10^{-5}$, 3.14, Lactic acid
 c. $1.38 * 10^{-4}$, 3.86, Lactic acid
 d. $9.17 * 10^{-4}$, 4.68, Formic acid

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4. What are the equilibrium values of carbonic acid and the pH of a 1.34M solution?

$[\text{H}_2\text{CO}_3] =$

$[\text{HCO}_3^-] =$

$[\text{CO}_3^{2-}] =$

pH =

5. Which salts yield neutral solutions?

- a. NH_4Cl
- b. CaCl_2
- c. LiNO_3
- d. $\text{Fe}(\text{NO}_3)_3$
- e. B and C
- f. A and D
- g. None of the above

6. Which of the following act as Lewis acids?

- a. Ba^{2+}
- b. NH_3
- c. AlCl_3
- d. H_2O
- e. A and C
- f. B and D

Chapter 19: Ionic Equilibria in Aqueous Systems

1. What is the pH of a buffer of 0.83M $(\text{CH}_3)_2\text{NH}_2\text{Cl}$ and 1.2M $(\text{CH}_3)_2\text{NH}$ before and after adding 125mL of 0.75M HCl to 1 L of the buffer. (Info: pK_b of $(\text{CH}_3)_2\text{NH}=3.23$).

- a. 8.776 -> 10.93
- b. 2.726 -> 7.901
- c. 10.93 -> 10.85
- d. 7.901 -> 2.726

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2. Magnesium phosphate is an anticaking agent for silicone-containing cleaning agents and salt. Its K_{sp} is 1.04×10^{-24} . If $[Mg^{2+}] = [PO_4^{3-}] = 3.6 \times 10^{-10}$ M, will magnesium phosphate precipitate?

- a. Yes, $Q_{sp} > K_{sp}$
- b. No, $Q_{sp} > K_{sp}$
- c. No, $Q_{sp} = K_{sp}$
- d. Yes, $Q_{sp} < K_{sp}$
- e. No, $Q_{sp} < K_{sp}$

3. Does the addition of HNO_3 affect the solubility of calcium fluoride?

- a. Increases solubility
- b. Decreases solubility
- c. No effect on solubility

4. What is the pH at the equivalence point of 912 mL of 10.67 M $HBrO$ with 15.02 M $NaOH$?

- a. 12.84
- b. 13.74
- c. 2.29
- d. 11.71
- e. 6.91